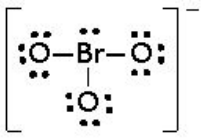
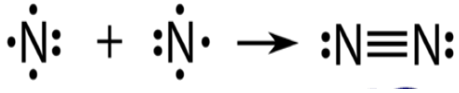
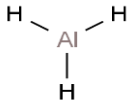
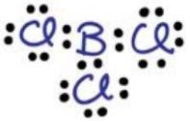
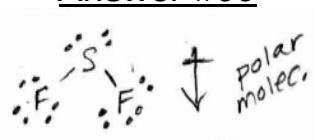
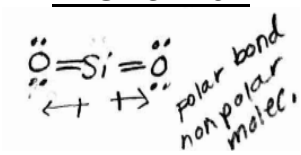
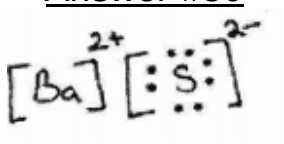
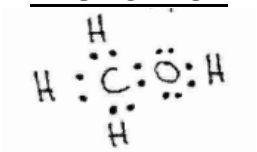
SET #1 **Questions #1 - 8**	**SET #2** **Questions #9 - 16**
Answer #1 **Strontium Oxide**	*Answer #9* **Single replacement/single displacement**
Answer #2 **421.61 g/mol**	*Answer #10* **Cl_2**
Answer #3 **Sea of electrons, delocalized electrons etc.**	*Answer #11* 
Answer #4 **Covalent bond**	*Answer #12* **tetrahedral**
Answer #5 **2 mol O_2 : 1 mol CH_4**	*Answer #13* **68.15 g/mol**
Answer #6 **Hydrogen bonding** **Remember H-NOF!**	*Answer #14* **trigonal pyramidal**
Answer #7 **Metallic < Ionic Lattice < Network Covalent**	*Answer #15* 
Answer#8 **$Cu(SO_4)_2$**	*Answer #16* **Combustion!**

SET #3 **Questions #17 - 25**
Answer #17 **$AlPO_4 + 3Rb(NO_2) \rightarrow Al(NO_2)_3 + Rb_3(PO_4)$**
Answer #18 **Unequally!**
Answer #19 **10 mol ZnO!** **Same molar ratio!**
Answer #20 **31.1g ZnO**
Answer #21 **H_2 F_2 Cl_2 Br_2** **I_2 O_2 N_2**
Answer #22 **77.98 g/mol**
Answer #23 **L of A→molA→mol B→L of B** **76L H_2O**
Answer #24 
Answer #25 **2 mole TNT : 7 mole CO**

SET #4 **Questions #26 - 33**
Answer #26 **$Na_2 CO_3$**
Answer #27 **106g/mol**
Answer #28 **$Fe_2(SO_4)_3$=** **400.1 g/mole**
Answer #29 **0.111 moles** **$Fe_2(SO_4)_3$**
Answer #30 **1087.7g** **KCl**
Answer #31 **1592.68 g** **$Fe_2(CO_3)_3$**
Answer #32 **Double displacement.**
Answer #33 **$3CuBr_2 + 2AlCl_3 \rightarrow 3CuCl_2 + 2AlBr_3$**

SET #5 **Questions #34 - 41**	**SET #6** **Questions #42 - 49**
Answer #34 **0.103 moles**	*Answer #42* **0.99g**
Answer #35 **1 mole $Fe_2(SO_4)_3$ =** **3 moles Na_2SO_4**	*Answer #43* **8.2 mol**
Answer #36 **30 moles Na_2SO_4**	*Answer #44* **1.23×10^{24} molecules**
Answer #37 **$2H_2 + O_2 \rightarrow 2H_2O$** **0.94 moles water**	*Answer #45* **5.44×10^{-5} mol B**
Answer #38 1. **Production of heat and light** 2. **Production of a gas** 3. **Formation of a precipitate** 4. **Change in color**	*Answer #46*  **Trigonal Planar**
Answer #39 **$2Na + Cl_2 \rightarrow 2NaCl$** **1 mole Cl_2**	*Answer #47* **Valence electrons**
Answer #40 **$2NaCl+Ba \rightarrow BaCl_2+2Na$** **39.32 g Na**	*Answer #48* **Gain 3 electrons**
Answer #41 **19.52L** **F_2**	*Answer #49* **26 ve-**

SET #7 Questions #50 - 58	SET #8 Questions #59 - 66
Answer #50 **Ionic, covalent, covalent, covalent, ionic**	*Answer #59* **1 lone pair**
Answer #51 **Mono, di, tri, tetra, penta, hexa, hepta, octa, nona, deca**	*Answer #60* 
Answer #52 **Dicarbon hexahhydride**	*Answer #61* 
Answer #53 **Silver oxide**	*Answer #62* **$CHCl_3$ is more polar because greater electroneg. difference between atoms than in $CHBr_3$**
Answer #54 **Copper (III) Nitrite**	*Answer #63* **London forces**
Answer #55 **Sulfur hexoxide**	*Answer #64* **Dipole-Dipole** **(yes H is there, but no H-NOF:)**
Answer #56 	*Answer #65* **Network Covalent**
Answer #57 	*Answer #66* **4**
Answer #58 **0 lone pairs**	